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Chemistry Organic Chemistry: A Guided Inquiry Use curved arrows to show the most likely acid base reaction between each boxed pair. a. Use the p K a values in Model 7 and in the previous ChemActivity to mark each curved arrow with a positive or negative energy change in p K a Units. b. Calculate ΔH for each reaction, and sketch an energy diagram showing ΔH as an arrow on this diagram.

Use curved arrows to show the most likely acid base ...

Chemistry Organic Chemistry: A Guided Inquiry The equilibrium constant for the reaction from CTQ 47 is given below. [Square brackets] represent the concentration of that ion or molecule in moles/liter at equilibrium. $K_{eq} = \frac{[NH_4^+][OH^-]}{[NH_3][H_2O]} = 10^{-7}$

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